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first law as  $w =$   
 $q - \Delta U$  ; thus  
 $q = \Delta H$ ;  $w = \Delta H - \Delta U$ ;  
4. Isothermal  
Process Because  
 $U$  is a function  
only of  $T$  for an

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ideal gas,  $DU = DH = 0$  for an isothermal process. These results also follow from the general results by using  $DT = D(PV) = 0$  for an isothermal process.

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. . . Determine

if the problem

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Problem 1.1\* The  
plot of  $V = V$

( $P, T$ ) for a gas  
is shown in Fig.

1.1. Determine.  
the expressions  
of the two

second  
derivatives of  
the volume of  
this plot.

(note: the  
principle

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curvatures of the surface are proportional to these second derivatives).

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calculate the  
values of  $T$ ,  $V$   
and  $P$  at the



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states A, B and  
C. Step 1 T V P  
A T V P B  
dynamics

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problems  
encountered in  
thermodynamics,  
the variables  
are limited to  
P, T, V,  $\epsilon_i$ ,  $s_i$ ,  
composition, and

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applied fields.  
The simplest  
examples involve  
only two  
variables. More  
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systems require  
more variables.  
8

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For most problems encountered in thermodynamics, the variables are limited to  $P$ ,  $T$ ,  $V$ ,  $\epsilon_i$ ,  $s_i$ , composition, and applied fields. The simplest examples involve only two variables. More complicated

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Systems require  
more variables.  
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encountered in  
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the variables  
are limited to  
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an object has a  
weight of 10 lbf  
on the moon,  
what would the  
same object

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