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2018 Nov HL paper 1 [IB Chemistry] - question-by-question  
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13.2 What are the Reasons for the Different Colours of Complexes?  
[HL IB Chemistry]

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RESULTS REACTION!! [May 2018 Session] | Katie Tracy~~

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Check A Level topical MCQ playlists - A Level Chemistry MCQ ☐  
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YOUR EXAM (FULL MARKS - 20/20)! | studycollab: Alicia ~~13.1~~

~~Transition Metals, their Complexes and Magnetism [HL IB  
Chemistry]~~ Understand the IB Chemistry SL + HL Assessment  
Format (Paper 1, 2, 3 and IA)

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Equilibrium HL paper 2 question: How to solve IB chemistry  
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point compared to oleic acid,  $\text{CH}_3(\text{CH}_2)_7\text{CH}=\text{CH}(\text{CH}_2)_7\text{COOH}$ .  
Outline your reason. [2]..... B3. (a) Some synthetic hormones are  
similar in structure to progesterone and estrogen and may be used to  
prevent pregnancy. Outline the mode of action of these hormones as  
oral contraceptives. [3].....

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HL Paper 3 Infrared (IR) spectroscopy is widely used as a technique in analytical chemistry. The IR spectrum, mass spectrum and spectrum of an unknown compound, Y, of molecular formula are as follows. (i) Identify the bonds responsible for the peaks A, B and C in the IR spectrum of Y. A: B: C:

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3. a ionic 1 3. b i red 1 3. b ii emission spectra of both  ${}^6\text{Li}$  and natural Li» give same colour/produce same «range of» wavelengths OR they have same electron transitions/same nuclear charge Accept [the spectra are almost identical]. 1 3. b iii ICP -MS Accept [MS/mass spectrometry]. 1 3. c r m 0.694 n 0.100 mol M 6.94 «

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(vi)  $\text{HCOOC}(\text{CH}_3)_3$  is an isomer of X. For the  $^1\text{H}$  NMR spectrum of this isomer, deduce the total number of peaks (excluding the TMS peak at 0 ppm) and the ratio of peak areas. For each peak, deduce whether it is a singlet, doublet, triplet, quartet or shows a more complex splitting pattern. [3] Number of peaks (excluding the TMS peak at 0 ppm):

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IB Chemistry SL and HL Core. Both IB Chemistry SL and HL cover the first 11 topics (a total of 95 hours), and HL additionally covers topics 12 to 21 (an additional 60 hours). For both levels, you'll also study one of the four options A-D (15 hours for SL, 25 hours for HL). Topic #1: Stoichiometric Relationships 13.5 Hours for SL and HL

The Complete IB Chemistry Syllabus: SL and HL  
volume made up to the graduated mark with distilled water. A 10.0 cm<sup>3</sup> sample of this new solution was pipetted into a new conical flask and titrated with 0.100 mol dm<sup>-3</sup> sodium hydroxide solution using phenolphthalein as the indicator. 12.60 cm<sup>3</sup> of the 0.100 mol dm<sup>-3</sup> sodium hydroxide solution was required to reach the end point.

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SL & HL Paper 3 Section A Experimental work Question 7

The 1:2 ratio of marks (paper 3-to-paper 2) for SL versus the 2:3 ratio for HL is important to bear in mind when deciding how much time to give over to each exam in preparation. On a similar note, the emphasis on the Option is different for SL and HL. The Option forms a much smaller fraction of the overall mark if you are studying it at SL.

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