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Question 1. Calculate the radius and
velocity of the electron in the n th orbit in
hydrogen atom from the following data, e

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 $= 1.6 \times 10^{-19} \text{ C}$, $m = 9.11 \times 10^{-31}$, $h = 6.625 \times 10^{-34} \text{ JS}$, $\lambda_0 = 8.85 \times 10^{-12} \text{ Fm}^{-1}$
and $c = 3 \times 10^8 \text{ s}^{-1} \dots$

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molecules, governed by the interaction of each atom or ... than the size of the nucleus (see Chapter 11, Section 11.6 in Class XI Physics textbook). Thus, the electrons would seem to be at a distance ... second law of motion and the Coulomb's law for electrostatic

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(a) Two stable isotopes of lithium have respective abundances of 7.5% and 92.5%. These isotopes have masses 6.01512 u and 7.01600 u, respectively. Find the atomic

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mass of lithium. (b) Boron has two stable isotopes, .

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susceptibility for (a) Diamagnetic substances (DMS) (b) Paramagnetic substances (PMS) (c) Ferromagnetic substances (FMS). Answer: (a) For DMS, χ is very low and negative $\approx 10^{-9}$ to -10^{-5} . e.g : For copper, χ is -9.8×10^{-6} .

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